

## Notes on the GSW function gsw\_SP\_salinometer

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This function calculates Practical Salinity (SP) from the conductivity ratio ( $R_t$ ) and in situ temperature ( $^{\circ}\text{C}$ , ITS-90) using the algorithm of the Practical Salinity Scale of 1978 (PSS-78, (Unesco (1981, 1983))). The conductivity ratio  $R_t$  is defined as (see Eqn. (E.2.1) of IOC *et al.* (2010))

$$R_t = \frac{C(S_p, t_{68}, 0)}{C(35, t_{68}, 0)},$$

being the ratio of two conductivity values, one from a general seawater sample and the other being a sample of Standard SeaWater (having a Practical Salinity of exactly 35), both being measured at the same value of temperature. A laboratory salinometer has this ratio of conductivities,  $R_t$ , as an output, and the present function, **gsw\_SP\_salinometer**, uses this conductivity ratio and the ITS-90 temperature  $t_{90}$  of the salinometer bath as the two input variables.

The first step is to calculate the  $t_{68}$  temperature from the  $t_{90}$  input bath temperature using Eqn. (A.1.3) of IOC *et al.* (2010), repeated here

$$(t_{68}/^{\circ}\text{C}) = 1.00024 (t_{90}/^{\circ}\text{C}). \quad (\text{A.1.3})$$

Practical Salinity  $S_p$  is then calculated from the PSS-78 expression, Eqn. (E.2.6) of IOC *et al.* (2010), repeated here

$$S_p = \sum_{i=0}^5 a_i (R_t)^{i/2} + \frac{(t_{68}/^{\circ}\text{C} - 15)}{[1 + k(t_{68}/^{\circ}\text{C} - 15)]} \sum_{i=0}^5 b_i (R_t)^{i/2}. \quad (\text{E.2.6})$$

Note that Practical Salinity is defined by PSS-78 only in the range  $2 < S_p < 42$ . For values of Practical Salinity less than 2, Millero *et al.* (2008) and IOC *et al.* (2010) outline the following method for estimating the Practical Salinity of a seawater sample. First Absolute Salinity is measured (via a direct measurement of density) and also Absolute Salinity Anomaly is estimated so that the Reference Salinity can be calculated from  $S_R = S_A - \delta S_A$ . Practical Salinity is then given by  $(35/35.165\ 04\ \text{g kg}^{-1}) S_R$ . For a seawater sample with a Practical Salinity greater than 42, the recommendation is to accurately dilute the sample with pure water and to then measure the diluted sample's conductivity ratio and use PSS-78 to determine its Practical Salinity. The Practical Salinity of the undiluted sample can then be calculated taking the known dilution factor into account.

In the absence of direct measurements of density, it is convenient to extend the range of applicability of **gsw\_salinometer** below a Practical Salinity of 2, and this is done by using the low-salinity extension of Hill *et al.* (1986). When the PSS-78 expression for Practical Salinity, Eqn. (E.2.6), gives a value less than 2, the **gsw\_salinometer** function returns a slightly modified version of the Hill *et al.* (1986) expression for Practical Salinity. The modification is necessary to ensure that the output of **gsw\_salinometer** is a continuous function of its inputs. Specifically, the Hill *et al.* (1986) algorithm expresses Practical Salinity in terms of  $R_t$  and  $t_{68}$ , being a correction to Eqn. (E.2.6) of IOC *et al.* (2010), when the Practical Salinity is less than 2. However, when the Practical Salinity is precisely 2 according to the PSS-78 expression Eqn. (E.2.6), the Hill *et al.* (1986) expression gives a Practical Salinity that is not exactly 2. This issue is overcome as follows. First the PSS-78 expression Eqn. (E.2.6) is used and if the resulting Practical Salinity is less than 2, a Newton-Raphson iterative technique (in the library function **gsw\_Hill\_ratio\_at\_SP2**) is used to find the value of  $R_t$  at which the Practical Salinity of PSS-78 is exactly 2 at the known  $t_{68}$  temperature. With these values of  $R_t$  and  $t_{68}$ , the Hill *et al.* (1986) expression is

evaluated, and the ratio of 2 to this value of Practical Salinity is noted. The final value of Practical Salinity from **gsw\_salinometer** is this ratio multiplied by the value of Practical Salinity derived from the straightforward use of the Hill *et al.* (1986) expression.

Given these definitions of Practical Salinity, it is clear that for a real sample of pure water or of seawater, Practical Salinity cannot be negative. Since salinity enters the TEOS-10 Gibbs function as powers of the square root of Absolute Salinity, even a slightly negative value of Absolute Salinity means that the outputs of many of the GSW functions become complex numbers. Hence it is convenient to ensure that the inputs to these functions do not encounter negative values of Absolute Salinity. One way this can be achieved is to ensure that Practical Salinity is never negative, and the present function **gsw\_salinometer** is constrained to never output a negative value of Practical Salinity.

### References

- Hill, K. D., T. M. Dauphinee and D. J. Woods, 1986: The extension of the Practical Salinity Scale 1978 to low salinities, *IEEE J. Oceanic Eng.*, **11**, 109–112.
- IOC, SCOR and IAPSO, 2010: *The international thermodynamic equation of seawater – 2010: Calculation and use of thermodynamic properties*. Intergovernmental Oceanographic Commission, Manuals and Guides No. 56, UNESCO (English), 196 pp. Available from <http://www.TEOS-10.org>
- Unesco, 1981: The Practical Salinity Scale 1978 and the International Equation of State of Seawater 1980. Unesco technical papers in marine science **36**, 25pp.
- Unesco, 1983: Algorithms for computation of fundamental properties of seawater. Unesco technical papers in marine science **44**, 53pp.

Here follows appendix E of the TEOS-10 Manual (IOC *et al.* (2010)).

## Appendix E: Algorithm for calculating Practical Salinity

### E.1 Calculation of Practical Salinity in terms of $K_{15}$

Practical salinity  $S_p$  is defined on the Practical Salinity Scale of 1978 (Unesco (1981, 1983)) in terms of the conductivity ratio  $K_{15}$  which is the electrical conductivity of the sample at temperature  $t_{68} = 15^\circ\text{C}$  and pressure equal to one standard atmosphere ( $p = 0$  dbar and absolute pressure  $P$  equal to 101 325 Pa), divided by the conductivity of a standard potassium chloride (KCl) solution at the same temperature and pressure. The mass fraction of KCl in the standard solution is  $32.4356 \times 10^{-3}$  (mass of KCl per mass of solution). When  $K_{15} = 1$ , the Practical Salinity  $S_p$  is by definition 35. Note that Practical Salinity is a unit-less quantity. Though sometimes convenient, it is technically incorrect to quote Practical Salinity in “psu”; rather it should be quoted as a certain Practical Salinity “on the Practical Salinity Scale PSS-78”. When  $K_{15}$  is not unity,  $S_p$  and  $K_{15}$  are related by (Unesco, 1981, 1983) the PSS-78 equation

$$S_p = \sum_{i=0}^5 a_i (K_{15})^{i/2} \quad \text{where} \quad K_{15} = \frac{C(S_p, t_{68} = 15^\circ\text{C}, 0)}{C(35, t_{68} = 15^\circ\text{C}, 0)}, \quad (\text{E.1.1})$$

and the coefficients  $a_i$  are given in the following table. Note that the sum of the six  $a_i$  coefficients is precisely 35, while the sum of the six  $b_i$  coefficients is precisely zero. Equation (E.1.1) is valid in the range  $2 < S_p < 42$ .

$i$	$a_i$	$b_i$	$c_i$	$d_i$	$e_i$
0	0.0080	0.0005	$6.766097 \times 10^{-1}$		
1	-0.1692	-0.0056	$2.00564 \times 10^{-2}$	$3.426 \times 10^{-2}$	$2.070 \times 10^{-5}$
2	25.3851	-0.0066	$1.104259 \times 10^{-4}$	$4.464 \times 10^{-4}$	$-6.370 \times 10^{-10}$
3	14.0941	-0.0375	$-6.9698 \times 10^{-7}$	$4.215 \times 10^{-1}$	$3.989 \times 10^{-15}$
4	-7.0261	0.0636	$1.0031 \times 10^{-9}$	$-3.107 \times 10^{-3}$	
5	2.7081	-0.0144			

### E.2 Calculation of Practical Salinity at oceanographic temperature and pressure

The following formulae from Unesco (1983) are valid over the range  $-2^\circ\text{C} \leq t \leq 35^\circ\text{C}$  and  $0 \leq p \leq 10\,000$  dbar. Measurements of salinity in the field generally measure the conductivity ratio  $R$

$$R = \frac{C(S_p, t_{68}, p)}{C(35, t_{68} = 15^\circ\text{C}, 0)} = \frac{C(S_p, t_{68}, p)}{C(S_p, t_{68}, 0)} \frac{C(S_p, t_{68}, 0)}{C(35, t_{68}, 0)} \frac{C(35, t_{68}, 0)}{C(35, t_{68} = 15^\circ\text{C}, 0)} \quad (\text{E.2.1})$$

which has been expressed in (E.2.1) as the product of three factors, which are labeled  $R_p$ ,  $R_t$  and  $r_t$  as follows

$$R = \frac{C(S_p, t_{68}, p)}{C(35, t_{68} = 15^\circ\text{C}, 0)} = R_p R_t r_t. \quad (\text{E.2.2})$$

The last factor  $r_t$  has been fitted to experimental data as the following polynomial in  $t_{68}$

$$r_t = \sum_{i=0}^4 c_i (t_{68} / ^\circ\text{C})^i \quad (\text{E.2.3})$$

and the factor  $R_p$  has been fitted to experimental data as a function of  $p$ ,  $t_{68}$  and  $R$  as

$$R_p = 1 + \frac{\sum_{i=1}^3 e_i p^i}{1 + d_1 (t_{68} / ^\circ\text{C}) + d_2 (t_{68} / ^\circ\text{C})^2 + R [d_3 + d_4 (t_{68} / ^\circ\text{C})]}. \quad (\text{E.2.4})$$

For any measurement of  $R$  it is possible to evaluate  $r_t$  and  $R_p$  and hence calculate

$$R_t = \frac{R}{R_p r_t}. \quad (\text{E.2.5})$$

At a temperature of  $t_{68} = 15^\circ\text{C}$ ,  $R_t$  is simply  $K_{15}$  and Practical Salinity  $S_p$  can be determined from (E.1.1). For temperatures other than  $t_{68} = 15^\circ\text{C}$ , Practical Salinity  $S_p$  is given by the following function of  $R_t$  with  $k = 0.0162$ ,

$$S_p = \sum_{i=0}^5 a_i (R_t)^{i/2} + \frac{(t_{68} / ^\circ\text{C} - 15)}{[1 + k(t_{68} / ^\circ\text{C} - 15)]} \sum_{i=0}^5 b_i (R_t)^{i/2}. \quad (\text{E.2.6})$$

Equations (E.1.1) and (E.2.6) are valid only in the range  $2 < S_p < 42$ . Outside this range  $S_p$  can be determined by dilution with pure water or evaporation of a seawater sample. Practical Salinity  $S_p$  can also be estimated using the extension of the Practical Salinity Scale proposed by Hill *et al.* (1986) for  $0 < S_p < 2$ . The GSW Oceanographic Toolbox incorporates a modified form of the extension of Hill *et al.* (1986) for  $0 < S_p < 2$ . The modification ensures that the algorithm is exactly PSS-78 for  $S_p \geq 2$  and is continuous at  $S_p = 2$ . The values of Practical Salinity  $S_p$  estimated in this manner may then be used in Eqn. (2.4.1), namely  $S_R \approx u_{PS} S_p$  to estimate Reference Salinity  $S_R$ .

When using a laboratory salinometer to evaluate Practical Salinity, use is made of Eqn. (E.2.6) since the salinometer returns  $R_t$  and the instrument's bath temperature is known (and is easily converted from a measured temperature on the ITS-90 scale to  $t_{68}$ ).

The temperatures in Eqns. (E.2.1) to (E.2.6) are all on the IPTS-68 scale. The functions and coefficients have not been refitted to ITS-90 temperatures. Therefore in order to calculate Practical Salinity from conductivity ratio at a measured pressure and  $t_{90}$  temperature, it is necessary first to convert the temperature to  $t_{68}$  using  $t_{68} = 1.00024 t_{90}$  as described Eqn. (A.1.3) of appendix A.1. This is done as the first line of the computer code described in the GSW Oceanographic Toolbox (appendix N). Further remarks on the implications of the different temperature scales on the definition and calculation of Practical Salinity can be found in appendix E.4 below.

### E.3 Calculation of conductivity ratio $R$ for a given Practical Salinity

When Practical Salinity is known and one wants to deduce the conductivity ratio  $R$  associated with this value of Practical Salinity at a given temperature, a Newton-Raphson iterative inversion of Eqn. (E.2.6) is first performed to evaluate  $R_t$ . Because  $r_t$  is a function only of temperature, at this stage both  $R_t$  and  $r_t$  are known so that Eqn. (E.2.4) can be written as a quadratic in  $R$  with known coefficients which is solved to yield  $R$ . This procedure is outlined in more detail in Unesco (1983). Computer software to perform this procedure is available in the GSW Oceanographic Toolbox as the functions **gsw\_R\_from\_SP** and **gsw\_C\_from\_SP** which return conductivity ratio and conductivity (in mS/cm) respectively. Note that this iterative inverse procedure is done in terms of  $t_{68}$ ; the code accepts  $t_{90}$  as the input and immediately converts this to a  $t_{68}$  temperature before performing the above iterative procedure.

## E.4 Evaluating Practical Salinity using ITS-90 temperatures

We first consider the consequence of the change from IPTS-68 to ITS-90 for the definition of Practical Salinity as a function of  $K_{15}$  and the defining mass fraction of KCl. Suppose Practical Salinity  $S_p$  were to be evaluated using the polynomial (E.1.1) but using  $K_{15-90}$  instead of  $K_{15}$ , where  $K_{15-90}$  is defined

$$K_{15-90} = \frac{C(S_p, t_{90} = 15 \text{ }^\circ\text{C}, 0)}{C(35, t_{90} = 15 \text{ }^\circ\text{C}, 0)}. \quad (\text{E.4.1})$$

The magnitude of the difference  $K_{15-90} - K_{15}$  can be calculated and is found to be less than  $6.8 \times 10^{-7}$  everywhere in the range  $2 < S_p < 42$ . Further calculation shows that  $\partial S_p / \partial K_{15} < 41$  everywhere in the valid range of Practical Salinity, so that the consequence of using  $K_{15-90}$  in (E.1.1) instead of  $K_{15}$  incurs a change in Practical Salinity of less than  $3 \times 10^{-5}$ . This is nearly two orders of magnitude below the measurement accuracy of a sample, and an order of magnitude smaller than the error caused by the uncertainty in the definition of the mass fraction of KCl. If all the original measurements that form the basis of the Practical Salinity Scale were converted to ITS-90, and the analysis repeated to determine the appropriate mass fraction to give the required conductivity at  $t_{90} = 15 \text{ }^\circ\text{C}$ , the same mass fraction  $32.4356 \times 10^{-3}$  would be derived.

Notwithstanding the insensitivity of this conductivity ratio to such a small temperature difference, following Millero *et al.* (2008a) the definition of Practical Salinity can be restated with reference to the ITS-90 scale by noting that the  $K_{15}$  ratio in Eqn. (E.1.1) can equivalently refer to a ratio of conductivities at  $t_{90} = 14.996 \text{ }^\circ\text{C}$ .

The fact that the conductivity ratio  $R_t$  is rather weakly dependent on the temperature at which the ratio is determined is important for the use of bench salinometers. It is important that samples and seawater standards be run at the same temperature, stable at order 1 mK, which is achieved by the use of a large water bath. However, it is not critical to know the stable bath temperature to any better than 10 or 20 mK.

The ratios  $R_p$ ,  $R_t$  and  $r_t$  that underlie the temperature-dependent expression (E.2.6) for Practical Salinity are more sensitive to the difference between IPTS-68 and ITS-90 temperatures and this is the reason why we recommend retaining the original computer algorithms for these ratios, and to simply convert the input temperature (which these days is on the ITS-90 temperature scale) in to the corresponding IPTS-68 temperature using  $t_{68} = 1.00024 t_{90}$  as the first operation in the software. Thereafter the software proceeds according to (E.2.1) – (E.2.6).

## E.5 Towards SI-traceability of the measurement procedure for Practical Salinity and Absolute Salinity

The observation of climate change taking place in the world ocean on a global scale over decades or centuries requires measurement techniques that permit the highest accuracy currently available, long-term stability and world-wide comparability of the measured values. The highest reliability for this purpose can be ensured only by traceability of these measurement results to the primary standards of the International System of Units (SI), supported by the National Metrological Institutes such as the NIST (National Institute of Standards and Technology) in the US, the NPL (National Physical Laboratory) in the UK, or the PTB (Physikalisch-Technische Bundesanstalt) in Germany.

In order to compute the thermodynamic properties of a seawater sample with standard composition, three independent parameters must be measured. Since the introduction of the Practical Salinity Scale of 1978 as an international standard for oceanography, these three properties have been electrolytic conductivity, temperature and pressure, from which salinity, density and other properties are computed in turn by

standard algorithms. The traceability of temperature and pressure measurement results, for example by CTD sensors, is ensured due to established calibration procedures carried out by the manufacturer or other laboratories and will not be considered here any further.

The observation of the ocean's salinity is a more complicated task (Millero *et al.* (2008a)). Even though over the last century different and permanently improved methods were developed and introduced in oceanography, traceability of salinity measurement results to SI units has not yet been achieved (Seitz *et al.* 2008 and Seitz *et al.* (2011)). This implies the risk that readings taken today may possess an enlarged uncertainty when being compared with observations taken a hundred years from now, a circumstance that will reduce the accuracy of long-term trend analyses performed in the future.

A quantity, quite generally, is a "property of a phenomenon, body or substance, where the property has a magnitude that can be expressed as a number" (ISO/IEC, 2007). The process to obtain this number is called measurement. The value of the indicated number (the quantity value) is determined by a calibration of the measuring system with a reference having a known quantity value of the same kind. In turn, the quantity value of the reference is assigned in a superior measurement procedure, which is likewise calibrated with a reference and so on. This calibration hierarchy ends in a primary reference procedure used to assign a quantity value and a unit to a primary standard for that kind of quantity. Thus, the unit of a measured quantity value expresses its link (its metrological traceability) to the quantity value of the corresponding primary standard. Obviously, quantity values measured at different times or locations, by different persons with different devices or methods can be compared with each other only if they are linked to the same reference standard, whose corresponding quantity value must be reproducible with a high degree of reliability.

Concerning comparability of measured quantity values a second aspect is of importance. The quantity value of a primary standard can only be realised with an inevitable uncertainty. The same holds for every measurement and calibration. A measurement result therefore always has to indicate the measured quantity value and its uncertainty. Obviously, the latter increases with every calibration step down the calibration hierarchy. Measured quantity values can evidently only be assumed equivalent if their difference is smaller than their measurement uncertainty (compatibility). On the other hand they can only be assumed reliably different, if the difference is larger than the uncertainty.

To ensure comparability in practice, the International System of Units (SI) was established. National Metrological Institutes (NMIs) have developed primary reference procedures to realise the SI units in the form of primary standards. Extensive (ongoing) efforts are made to link these units to fundamental and physical constants in order to achieve the highest degree of reproducibility. Moreover, the NMIs periodically conduct international comparison measurements under the umbrella of the International Bureau of Weights and Measures, in order to ensure the compatibility of the quantity values of national standards.

PSS-78, and similarly the Reference-Composition Salinity Scale (Millero *et al.* (2008a)), compute the salinity value from a measured conductivity ratio with respect to the  $K_{15}$  conductivity ratio of IAPSO Standard Seawater (SSW, Culkin and Ridout (1998), Bacon *et al.* (2007)), which plays the role of a primary standard. The production procedure of IAPSO Standard Seawater, and in particular the adjustment of its conductivity to that of a potassium chloride (KCl) solution of definite purity and the corresponding assignment of the  $K_{15}$  ratio, can be seen as a primary reference procedure. However both of these solutions are artefacts lying outside the SI system; they are not subject to regular international inter-comparisons; their sufficiently precise replicability by arbitrary independent laboratories is neither known nor even granted. A slow drift of artefact

properties cannot rigorously be excluded, similar in principle to the “evaporation” of mass from the kilogram prototype stored in Paris. It is impossible to foresee effects that might affect the conductivity of SSW solution one day. Thus, with respect to decadal or century time scales, there is an uncertainty of its  $K_{15}$  ratio, which a priori cannot be quantified and puts long term comparability of salinity measurement results at risk.

This fundamental problem, which is related to any artificial reference standard, can, at least in principle, be avoided if the conductivity of seawater is measured traceable to primary SI standards (“absolute” conductivity) rather than relying on a conductivity ratio. Unfortunately the related uncertainty of absolute conductivity measurements with present-day state-of-the-art technology is one order of magnitude larger than that of the relative measurements presently used in oceanography (Seitz *et al.* (2008)).

A way out of this practical dilemma is the measurement of a different seawater quantity that is traceable to SI standards and possesses the demanded small uncertainty, and from which the salinity can be computed via an empirical relation that is very precisely known (Seitz *et al.* (2011)). Among the potential candidates for this purpose are the sound speed, the refractive index, chemical analysis (e.g. by mass spectroscopy) of the sea-salt constituents, in particular chlorine, and direct density measurements. The latter has three important advantages, i) SI-traceable density measurements of seawater can be carried out with a relative uncertainty of 1 ppm (Wolf (2008)), which perfectly meets the needs of ocean observation, ii) a relation exists between density and the Absolute Salinity of seawater is available with a relative uncertainty of 4 ppm in the form of the TEOS-10 Gibbs function, iii) the measurand, density, is of immediate relevance for oceanography, in contrast to other options.

It is important to note that the actual measuring procedure for a quantity value is irrelevant for its traceability. To measure the weight of a person, a mass balance can be used, a spring or a magnetic coil; it is the quantity value that is traceable, not the method to achieve this value. The method in use is not intrinsically important except in so far as it is responsible for the uncertainty of the quantity value. Hence, we may measure the density of seawater with a CTD conductivity sensor, provided this sensor is properly calibrated with respect to an SI-traceable density reference standard. In practice, this will mean that the sensor calibration in oceanographic labs must be done with standard seawater samples of certified density rather than certified Practical Salinity. The density value returned from the CTD reading at sea is then converted into an Absolute Salinity value by means of the equation of state of seawater, and eventually into a Practical Salinity number for storage in data centres. The latter step may include some modification regarding local sea salt composition anomalies. Storing a salinity value rather than the related density reading has the advantage of conservativity with respect to dilution or changes of temperature or pressure.

This conceptual proposal of WG127 is still immature and needs to be worked out in more detail in the following years. Although it may imply only minor changes in the practical use of a CTD or similar devices, the new concept is very promising regarding the long-term reliability of observations made in the near future for climatic trend analyses to be performed by the coming generations. An immediate consequence of this proposal is to have the density (at a given temperature and pressure) of several samples of each batch of IAPSO Standard Seawater measured when they are produced and have these densities made available as reference values for each batch.